



Name \_\_\_\_\_

EOR 1  
Chapter 2 Reading Guide

2.1

- A) List the 4 most common elements in your body?
- B) List 7 other trace elements that are also essential for your living body?

2.2

- A) Explain why most table salt has a little bit of Iodine added to it?

2.3

- A) Explain why salt is a **compound** ?
- B) Why is salt a great example of matter having emergent properties?

2.4

- A) Explain the difference between **atomic number** and **mass number** for an atom?
- B) Define **isotope** and give an example?

2.5

- A) Explain how an isotope could **HELP** a human?
- B) Explain how an isotope could **HARM** a human?

2.6

- A) Atoms are most stable and **least likely to react** with other atoms when they have \_\_\_\_\_

2.7

- A) What would cause a neutral atom to become a positively charged **ion**?
- B) Explain what holds two atoms together that form an **ionic bond**?

2.8 & 2.9

- A) What do two atoms do with their electrons when they form a **nonpolar** covalent bond?
- B) What do two atoms do with their electrons when they form a **polar** covalent bond?

2.10

- A) What allows neighboring water molecules to form **hydrogen bonds**?

2.11

A) Define **cohesion** and explain how cohesion is used by a tree for daily survival?

2.12

A) Explain why **water resists temperature change** more than most other substances?

B) Explain how this property **benefits** life on earth?

2.13

A) Explain why solid ice is **less dense** than liquid water?

2.14

A) Explain what determines if a substance will or won't dissolve in water?

2.15

A) Draw a diagram of the pH scale below and identify where 10 or more different liquids would belong on the scale

B) Compare the definitions of an acid and a base?

2.18

A) Compare the definitions of reactants and products.

B) Draw an example chemical reaction and circle the products